

**RESEARCH ARTICLE**

# Heat flow inside a catalyst particle for steam methane reforming: CFD-modeling and analytical solution

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## Summary

Numerical investigation of a steam methane reforming process was performed from point of view to understand the heat flows inside a catalyst particle. To verify the numerical results, a new method based on the thermal conductivity equation has been developed to determine the temperature distribution inside the catalyst particle. The CFD-model was realized via ANSYS Fluent. To model the steam methane reforming process, the industrial Ni-based catalyst with a spherical particle was chosen. The temperature contours inside the catalyst particle and hydrogen mole fraction in the reaction space was calculated both numerically and analytically. The results show the irregularity in the distribution of the temperature field inside the catalyst. In the direction of flow, a minimum catalyst temperature occurs. In this case, the temperature decrease inside the catalyst occurs unevenly. Also, the temperature change on the catalyst surface as a function of flow time was analyzed.

## KEYWORDS:

CFD-modeling, steam methane reforming, thermal conductivity, temperature distribution

## 1 | INTRODUCTION

Computational fluid dynamics modeling (CFD-modeling) of various physical and chemical processes has recently been widely used by engineers and scientists in many countries. The different special software is used for CFD-modeling. The next software for CFD-modeling as ANSYS<sup>1,2</sup>, Comsol<sup>3</sup>, Autodesk Simulation CFD<sup>4</sup>, OpenFOAM (open source) are most widely used. Also, the constant increase in the computing power of personal computers makes it possible to use CFD-modeling for a wide range of the researcher's tasks.

Among the problems that can be solved using CFD-modeling, an important place is occupied by CFD-modeling of the methane steam reforming (SMR) process<sup>5,6,7</sup>. In recent years, notable successes have been achieved in the numerical study of

the SMR process. There are a large number of publications that discuss various aspects of the numerical simulation of SMR: the effect of technological parameters, the effect of geometric dimensionality of the computational domain, local transport and reaction rates, etc. The importance of CFD modeling of SMR is determined primarily by the wide application of this process in the industry: chemistry, energy, transport, etc. Tran et. al presented the results of CFD modeling of the SMR process in the industrial-scale steam methane reforming furnaces<sup>8</sup>. The authors concluded that a CFD-model of a steam methane reformer can be considered to be an adequate representation of the on-line reformer and can be used to determine the risk to operate the on-line reformer at un-explored and potentially more beneficial operating conditions. In addition, Lao et al. published the results of numerical investigation and control of the industrial steam methane reforming reactor via CFD-modeling<sup>9</sup>. They numerically investigated the industrial-scale top-fired co-current reformer that heated by natural gas combustion.

Most of the works dealing with CFD-modeling of SMR investigate the distribution of temperature, velocity, species concentration, etc. in the reaction zone. The results of these simulations can be verified using experimental data, which can be obtained using relatively simple laboratory facilities. Despite the notable successes achieved in the CFD-modeling of methane steam reforming, there are still a large number of practically interesting problems that need to be solved. One of these tasks is to study the temperature distribution inside a catalyst particle when SMR occurs. The complexity of this task lies in the fact that the experimental results inside the catalyst particles are rather difficult to obtain. However, experimental data are an important component for verifying the CFD-modeling results.

In this study, CFD-modeling of the steam methane reforming process over a pre-heated Ni-based catalyst was performed to determine the temperature distribution inside the catalyst. Moreover, to verify the CFD-modeling results, a new method has been developed to determine the temperature distribution inside the catalyst particle, based on the thermal conductivity equation. The objectives of the study are:

- CFD modeling of the SMR processes and determination of temperature distribution in the catalyst;
- Development of a new method for the analytical determination of the temperature distribution in the catalyst.

## 2 | CFD-MODELING

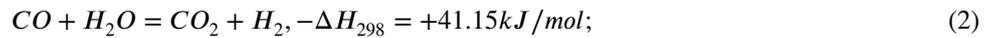
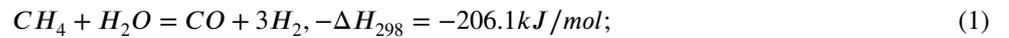
### 2.1 | Computational domain and mesh

Typically, the steam methane reforming process is carried out in the reformers with a fixed bed, which is filled with the catalyst particles. The shapes of the catalyst are different: cylinders, spheres, Raschig rings, etc. The effect of the catalyst shapes on the reforming characteristics have been reported in the authors' previous papers<sup>10</sup>. One of the main advantage of CFD-modeling is the fact that catalysts of various shapes can be used for modeling.

In this study, the computational domain is selected taking into account the further prospects of verification using numerical methods. Therefore, the spherical catalyst is selected for the computational domain. In addition, several assumptions and simplifications were made in the computational domain: the reaction space contains only one catalyst particle 40 mm in diameter; the reformer wall is adiabatic; the catalyst temperature and the temperature of the initial reaction mixture are the same at the initial time.

A flow region with no reaction is added before the catalyst particle to obtain a developed flow near a catalyst particle. The schematic view of the computational domain is presented in Fig.1.

SMR is widely investigated process. In the reaction space of the steam methane reformer various chemical reactions occurs. But there are tree main reactions in the temperature range above 800K<sup>11</sup>: steam methane reforming reactions (1) and (3), and water-gas shift reaction (WGSR) (2):



In the computational domain, the chemical reactions occur on the catalyst. Heat for the reaction is supplied by reaction flow. The input heat is consumed during endothermic reactions of the SMR process. There is no heat flow through the wall. Such a process corresponds to methane reforming over a pre-heated catalyst<sup>?</sup>.

The length of the reaction space is 200 mm, the diameter is 80 mm. The physical property catalyst is shown in Table 1.

The mesh for the computational domain in Fig.1 contains about 2 million elements. The inflation layers around the solid body (catalyst) were insert for the flow domain. These inflation layers allows to obtaint more accurate mesh structure. Moreover, the inflation layers are also added near the tube wall in the fluid domain. The total thickness of the inflation layer is calculated to get  $y^+$  such as  $20 < y^+ < 250$ .

## 2.2 | Governing equations

The CFD-model of the steam methane reforming process was realized via ANSYS Fluent on IntelXeon E3-1690 server with double 20-core processors. In the CFD-model, the conservation equations of mass and momentum, as well as the energy equation in combination with the species transport equations were solved in ANSYS Fluent 18.2. The transient (unsteady) flow was modeled. To simplify the model, the following assumptions are taken for the model: no the flux of energy due to a mass concentration

gradient (no Dufour effects)<sup>?</sup>; work by viscous forces and by pressure is not done; surface oxidation reactions of the metal wall are absent.

The mass conservation equation for unsteady flow have the following form:

$$\frac{\partial \rho}{\partial t} + \nabla \cdot (\rho \vec{v}) = 0 \quad (4)$$

The equation of conservation of momentum can be written as follows<sup>?</sup>:

$$\frac{\partial}{\partial t}(\rho \vec{v}) + \nabla \cdot (\rho \vec{v} \vec{v}) = -\nabla \cdot p + \nabla \cdot (\bar{\bar{\tau}}) + \rho \vec{g} + \vec{F} \quad (5)$$

where  $p$  is the static pressure;  $\rho \vec{g}$  and  $\vec{F}$  are the gravitational body force and external body forces, respectively. The stress tensor  $\bar{\bar{\tau}}$  is given by follows equation:

$$\bar{\bar{\tau}} = \mu \left[ (\nabla \vec{v} + \vec{v}^T) - \frac{2}{3} \nabla \cdot \vec{v} I \right] \quad (6)$$

where  $\mu$  is the molecular viscosity,  $I$  is the unit tensor, and the second term on the right hand side is the effect of volume dilation.

The energy conservation equations for each reaction component in the reaction zone are expressed as:

$$\nabla \cdot \vec{u}(p + \rho \cdot h_f) = \nabla \cdot \left[ k_{eff} \cdot \nabla T - \left( \sum h_j \cdot \vec{J}_j \right) + \bar{\bar{\tau}} \cdot \vec{v} \right] + S_f^h \quad (7)$$

In eq. (7)  $T$  is temperature of mixture ;  $\rho$  is the density of the  $j$ th component;  $h_f$  is enthalpy of the reaction mixture;  $\vec{J}_j$  is the diffusion flux of  $j$  element;  $k_{eff}$  is the conductivity;  $h_j$  is the enthalpy of  $j$  element;  $S_f^h$  is the fluid enthalpy source term (include the source of energy due to chemical reaction).

The energy conservation equation for the steam methane reformer walls has the following form:

$$\nabla \cdot (k_w \cdot \nabla T) = 0 \quad (8)$$

where  $k_w$  is the thermal conductivity of a wall.

The transport equations of  $j$  element in the reaction mixture can be written as follows:

$$\nabla \cdot (\rho \cdot \vec{v} \cdot Y_j) = -\nabla \cdot \vec{J}_j + R_j \quad (9)$$

where  $R_j$  is the chemical reaction rate of formation or decomposition of  $j$ th element;  $Y_j$  is the mass fraction of  $j$  element in the reaction mixture;  $\vec{J}_j$  is diffusion flux of  $j$  element.

The equation for diffusion flux of  $j$  element is given by follows equation:

$$\vec{J}_j = - \left( \rho \cdot D_{j,m} + \frac{\mu_t}{Sc_t} \nabla Y_j - D_{T,j} \frac{\nabla T}{T} \right) \quad (10)$$

where  $D_{j,m}$  and  $D_{T,j}$  is mass diffusion coefficient and thermal diffusion coefficient for  $j$  element, respectively;  $Sc_t$  and  $\mu_t$  are the turbulent Schmidt number and turbulent viscosity, respectively.

## 2.3 | Turbulence model

The method of Reynolds-averaged Navier-Stokes equations (RANS) was used to simulate the steam methane reforming process.

To solve RANS the RNG  $k$ - $\epsilon$  model is chosen<sup>??</sup>.

$$\frac{\partial}{\partial x_i}(\rho k v_i) = \frac{\partial}{\partial x_j} \left( \alpha_k \mu_{eff} \frac{\partial k}{\partial x_j} \right) + G_k + G_b - \rho \epsilon - Y_M + S_k \quad (11)$$

and

$$\frac{\partial}{\partial x_i}(\rho \epsilon v_i) = \frac{\partial}{\partial x_j} \left( \alpha_\epsilon \mu_{eff} \frac{\partial \epsilon}{\partial x_j} \right) + C_{1\epsilon} \frac{\epsilon}{k} (G_k + C_{3\epsilon} G_b) - C_{2\epsilon} \rho \frac{\epsilon^2}{k} - R_\epsilon + S_\epsilon \quad (12)$$

In eq. (11) and (12) the following notation are used:

- $G_k$  – the generation of turbulence kinetic energy due to the mean velocity gradients;
- $G_b$  – the generation of turbulence kinetic energy due to buoyancy;
- $Y_M$  – the contribution of the fluctuating dilatation in compressible turbulence to the overall dissipation rate;
- $\alpha_k$  and  $\alpha_\epsilon$  – the inverse effective Prandtl numbers for  $k$  and  $\epsilon$ , respectively;
- $S_k$  and  $S_\epsilon$  – user-defined source terms.

$R_\epsilon$  source in the eq.(12) is given by follows equation:

$$R_\epsilon = \frac{C_\mu \rho \eta^3 (1 - \eta/\eta_0) \epsilon^2}{1 + \beta \eta^3} \frac{1}{k} \quad (13)$$

where  $\eta \equiv S k / \epsilon$ ,  $\eta_0 = 4.38$ ,  $\beta = 0.012$ .

The constants  $C_{1\epsilon}$  and  $C_{2\epsilon}$  in  $k$ - $\epsilon$  turbulence model have the following values<sup>?</sup>:

$$C_{1\epsilon} = 1.42; \quad C_{2\epsilon} = 1.68 \quad (14)$$

## 2.4 | Kinetic model

In this model the reaction rates for the reaction element were determined based on the recognized research performed by Froment and Xu<sup>11</sup>. They suggested that both methane and steam is adsorbed on the surface with dissociation.  $\text{CH}_4$  consumption reaction rate for the reactions (1 and 3) is determined as:

$$r_{\text{CH}_4} = R_1 + R_3. \quad (15)$$

The reaction rate of formation of carbon dioxide ( $\text{CO}_2$ ) and carbon monoxide (CO) is determined as:

$$r_{\text{CO}_2} = R_3 + R_2. \quad (16)$$

$$r_{CO} = R_1 - R_2. \quad (17)$$

The expressions for determination of the rate of the reactions (1)-(3) have the following forms<sup>12</sup>:

$$R_1 = \frac{k_1}{p_{H_2}^{2.5}} \left( p_{CH_4} \cdot p_{H_2O} - \frac{p_{H_2}^3 \cdot p_{CO}}{K_{e1}} \right) \cdot \frac{1}{Q_r^2}; \quad (18)$$

$$R_2 = \frac{k_2}{p_{H_2}} \left( p_{CO} \cdot p_{H_2O} - \frac{p_{H_2} \cdot p_{CO_2}}{K_{e2}} \right) \cdot \frac{1}{Q_r^2}; \quad (19)$$

$$R_3 = \frac{k_3}{p_{H_2}^{3.5}} \left( p_{CH_4} \cdot p_{H_2O}^2 - \frac{p_{H_2}^4 \cdot p_{CO_2}}{K_{e3}} \right) \cdot \frac{1}{Q_r^2}; \quad (20)$$

$$Q_r = 1 + K_{CO} \cdot p_{CO} + K_{H_2} \cdot p_{H_2} + K_{CH_4} \cdot p_{CH_4} + \frac{K_{H_2O} \cdot p_{H_2O}}{p_{H_2}}, \quad (21)$$

where  $p_i$  is the partial pressure of  $j$ th reaction element, bar;  $K_{j_e}$  is the equilibrium constants of 1, 2, 3 reactions, respectively;  $k_j$  is kinetic rate constant of  $j$ th reaction.

In eq.(21) adsorption constants of  $i$  element ( $K_i$ ) have the following form:

$$K_i = K_{0i} \cdot e^{-\Delta H_i / (R_0 T)}, \quad (22)$$

where  $H_i$  is the adsorption enthalpy of  $i$  element, J/mol;  $K_{0i}$  is constant of  $i$  component.

## 2.5 | Boundary conditions

In this CFD-model the pressure-based solver is chosen. The convergence criterion for the mass and momentum equations, and species transport equations are set as  $10^{-3}$ , for the energy equation the convergence criterion is  $10^{-6}$ . The hybrid initialization (HI) is done to initialize model setup. The Second-Order Upwind approximation method is chosen to get more accurate results. The initial conditions for the model are presented in Table 2.

Also the boundary conditions for the reformer geometry presented in Fig.1 are set as follows:

– a reformer inlet:

$$x = 0; T_g = T_{g(in)}; C_i = C_{i(in)}; T_g = T_{cat}. \quad (23)$$

– a reformer outlet:

$$x = L; \frac{\partial C_i}{\partial x} = 0. \quad (24)$$

– a reformer center:

$$y = 0; \frac{\partial C_i}{\partial y} = 0; \frac{\partial T_g}{\partial y} = 0. \quad (25)$$

where  $C_i$  is mole concentration of  $i$  element.

### 3 | RESULTS AND DISCUSSION

#### 3.1 | Analytical solution

To verify the temperature and heat flux distribution model inside the catalyst particle, a new algorithm was developed. Such an algorithm allows us to obtain the exact distribution of temperature and heat flux inside the catalyst based on the fundamental equations of heat conduction.

The heat transfer process in the catalyst can be described by the parabolic heat equation. In the developed model, it is assumed that the thermophysical properties of the catalyst are isotropic and constant over time, therefore, the equation has the following form:

$$\frac{\partial T}{\partial t} = a \nabla^2 T + \frac{q_{smr}}{C_v}, \quad (26)$$

where  $T$  - temperature;  $t$  - time;  $a = k/C_v$  - thermal diffusivity coefficient;  $k$  - thermal conductivity;  $C_v = C_p \cdot \rho$  - specific volumetric heat capacity;  $\rho$  - density;  $q_{smr}$  - endothermic effect of the steam methane reforming process.

The last term of eq.(26) takes into account the absorption of heat due to the steam methane reforming reaction. To obtain a solution to equation (26), the boundary-value problem describing the temperature distribution in a spherical catalyst was considered. At the initial time, the temperature of the catalyst is equal to the temperature of the reaction mixture. On the surface of the catalyst, a boundary condition of the third kind is set with a constant temperature of the reaction mixture  $T_g$  and the heat transfer coefficient  $h$ . On the surface of the catalyst, there are uniformly distributed heat sinks caused by the endothermic reaction of the steam methane reforming process.

Given the isotropy of catalyst thermophysical properties, the uniform distribution of heat sinks and heat fluxes on the catalyst surface, as well as the constancy of the initial temperature at each point of the studied region, the temperature will be a function of two variables - spatial ( $r$ ) and temporal ( $t$ ). Therefore, eq.(26) can be written as:

$$\frac{\partial T(r; t)}{\partial t} = a \left[ \frac{2}{r} \frac{\partial T(r; t)}{\partial r} + \frac{\partial^2 T(r; t)}{\partial r^2} \right] + \frac{q_{smr}}{C_v}. \quad (27)$$

The boundary conditions for eq.(27) have the following form:

$$T(r; 0) = T_0; \left. \frac{\partial T(r; t)}{\partial r} \right|_{r=0} = 0; -k \left. \frac{\partial T(r; t)}{\partial r} \right|_{r=R} = h(T(R; t) - T_g). \quad (28)$$

where  $h$  - heat-exchange coefficient;  $R$  - radius of the catalyst particle;  $T_g$  - reaction mixture temperature.

The solution of eq.(27) can be obtained by applying the Laplace integral transform. Using the transition tables to the originals of functions, the solution of eq.(27) has the following form:

$$T(r; t) = T_g + \frac{q_{smr} R^2}{6k} \left( 1 + \frac{2kR}{h} - \frac{r^2}{R^2} \right) - \sum_{n=1}^{\infty} \left( T_g - T_0 + \frac{q_{smr} R^2}{k\mu_n^2} \right) A_n \frac{R \sin\left(\mu_n \frac{r}{R}\right)}{r\mu_n} \exp\left(-\mu_n^2 \frac{at}{R^2}\right). \quad (29)$$

The expression (29) is obtained for the case when the temperature of the catalyst is not equal to the temperature of the reaction mixture.

Using the approximate method<sup>13</sup> for solving equations (27)-(28), based on the introduction of a new function and additional boundary characteristics, a solution can also be found in the following form:

$$\Theta(\xi; Fo) = \frac{Po}{6} \left( \xi^2 - 1 - \frac{2}{Bi} \right) \exp(-\eta Fo) + \frac{Po}{6} \left( 1 - \xi^2 + \frac{2}{Bi} \right) + 1. \quad (30)$$

where  $\Theta(\xi; Fo) = \frac{T(r;t)}{T_0}$  - dimensionless temperature;  $Po = \frac{q_{smr} R^2}{kT_0}$  - Pomerantsev criterion;  $\eta = \frac{9Bi}{Bi+3}$ ;  $Bi = \frac{h}{kR}$  - Biot number;  $\xi$  - relative coordinate;  $Fo = \frac{at}{R^2}$  - dimensionless time;  $\xi = \frac{r}{R}$  - relative distance.

The enthalpy of the endothermic reaction was determined on the kinetic model suggested by Xu and Froment. The kinetic model was discussed above 2.4.

The equation (30) can be used in the case when the heat consumption is uniformly distributed over the entire body volume (according to the last term of equation (27)). If the source proceed only on the catalyst surface, boundary conditions for eq.(27) have the following form:

$$T(r; 0) = T_0; \frac{\partial T(r; t)}{\partial r} = 0; -k \frac{\partial T(R; t)}{\partial r} = q_{smr} - h(T_g - T(R; t)). \quad (31)$$

The equation for determination of temperature distribution in the Catalyst particle will be written as follows:

$$\Theta(\xi; Fo) = \frac{B}{T_0} \left[ \frac{3 \exp(-\eta \cdot Fo) (T_0 B^{-1} - 1) Bi (1 + 2Bi^{-1} - \xi^2)}{2Bi + 3} \right] + 1 \quad (32)$$

where  $B = T_g - \frac{q_{smr}}{h}$ .

The equation (32) has great practical importance. Using a simple expression (32), it is possible to determine the temperature distribution in any catalyst whose thermal properties are known for any endothermic reaction. This suggests that the equation (32) can be used not only to model heat fluxes inside the catalyst for the steam reforming reaction but for other endothermic reactions.

### 3.2 | Contours and profiles

The results of CFD modeling are presented as the contours and profiles of various controlled parameters of the steam methane reforming process. The main advantage of the results of CFD modeling is the visibility of the results.

Fig.2 shows the temperature profiles inside the catalyst and the hydrogen mole fraction for a time on stream of 5 seconds. The initial data for the obtained results are as follows: the initial temperature of the catalyst is 1000K, the composition of the initial reaction mixture is  $H_2O:CH_4:H_2 = 2:1:0.15$ , and the flow rate of the reaction mixture is 1 m/s.

The results in Fig.2 are obtained when solving a non-stationary problem for a time moment of 5 seconds after the start of the flow. The CFD model was obtained via ANSYS Fluent 18.2 Academic version. Convergence for each numerical experiment took approximately 6000 iterations for which the computation time was on the order of 30 hours.

Fig.2 clearly shows the irregularity in the distribution of the temperature field inside the catalyst. In the direction of flow, a minimum catalyst temperature occurs. In this case, the temperature decrease inside the catalyst occurs unevenly. The maximum temperature of about 993 K does not occur in the center of the particle but is shifted in the direction opposite to the motion of the catalyst particles. Such an irregularity in the temperature change can be explained by the fact that in the frontal part of the catalyst the intensity of the methane steam reforming reaction is maximum. Moreover, in the back of the reaction intensity decreases.

In addition, Fig.2 shows the contours of the mole fraction of hydrogen. Fig.2 depicts that when passing the reaction mixture near the surface of the catalyst, the hydrogen concentration increases. This fact correlates well with the works of other authors.

To quantitatively compare the results of CFD modeling and numerical solutions, the temperature profile at various points in time was analyzed.

Fig.3 shows a temperature profile inside the catalyst. The solid line is the results of the analytical solution, the dashed line is the results of CFD modeling. In order to obtain the dependence of temperature on the radial position for CFD modeling, auxiliary circles were constructed. Therefore, in fig. Fig.3 shows the average temperature for each radial position within the catalyst.

Fig.3 shows that at the initial time, the temperature inside the catalyst is constant. Then, when the reaction mixture gets on the catalyst, the endothermic reaction of methane steam reforming begins to proceed. This reaction proceeds mainly on the surface of the catalyst; therefore, the temperature on the surface of the catalyst is lower than inside the catalyst. With an increase in the flow time, the temperature profile gradually becomes equal, and approximately 25 seconds after the start of the flow, the temperature inside the catalyst and on the surface becomes equal. With the help of such a study, an important conclusion can be made, namely, for the studied initial data, approximately 25 seconds after the start of the heat flux that contains the reaction mixture, it will be enough to compensate for the endothermic effect of the methane steam reforming reaction.

Fig.3 shows the temperature change on the catalyst surface as a function of flow time. Also, Fig.3 shows auxiliary circles for obtaining the results of CFD modeling. Fig.3 makes it clear that at a flow time of about 25 seconds, the temperature on the surface of the catalyst stabilizes.

## 4 | CONCLUSION

It has been shown that a commercial CFD code can be used to determine the heat flow inside a catalyst particle. To verify the results of CFD-modeling, the new method based on the thermal conductivity equation was developed. The new equation has great practical importance. Using a simple expression, it is possible to determine the temperature distribution in any catalyst whose thermal properties are known for any endothermic reaction. This suggests that the equation can be used not only to model heat fluxes inside the catalyst for the steam reforming reaction but for other endothermic reactions. The discrepancy between the results obtained via CFD-model and via new method was less 5%. Based on the CFD-model it was established the irregularity in the distribution of the temperature field inside the catalyst. In the direction of flow, a minimum catalyst temperature occurs. Also, the temperature change on the catalyst surface as a function of flow time was analyzed. It was established that at a time on stream of about 25 seconds, the temperature on the surface of the catalyst stabilizes.

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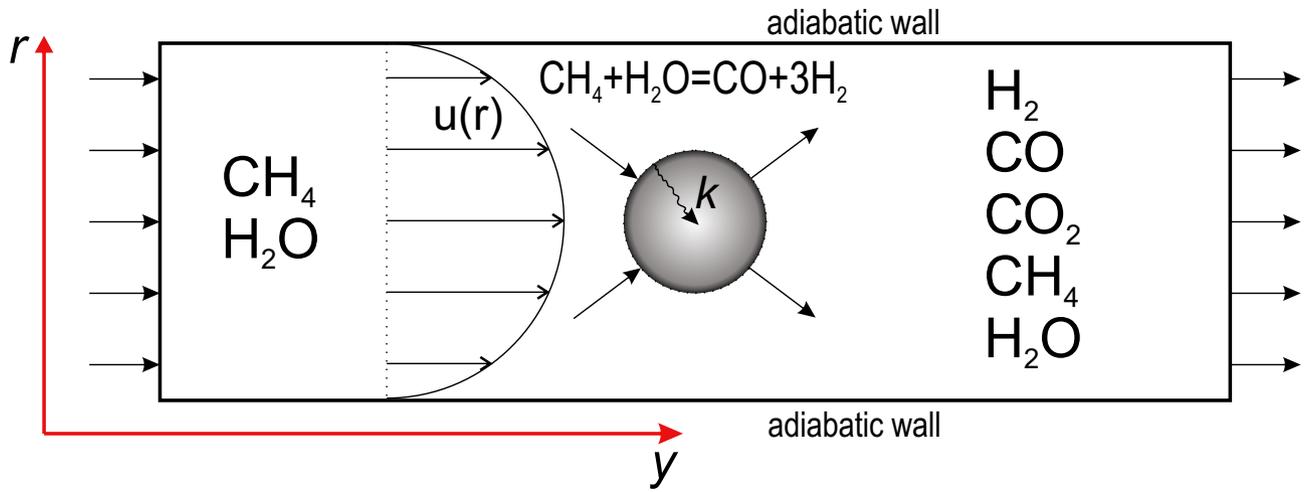
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**TABLE 1** Catalyst properties.

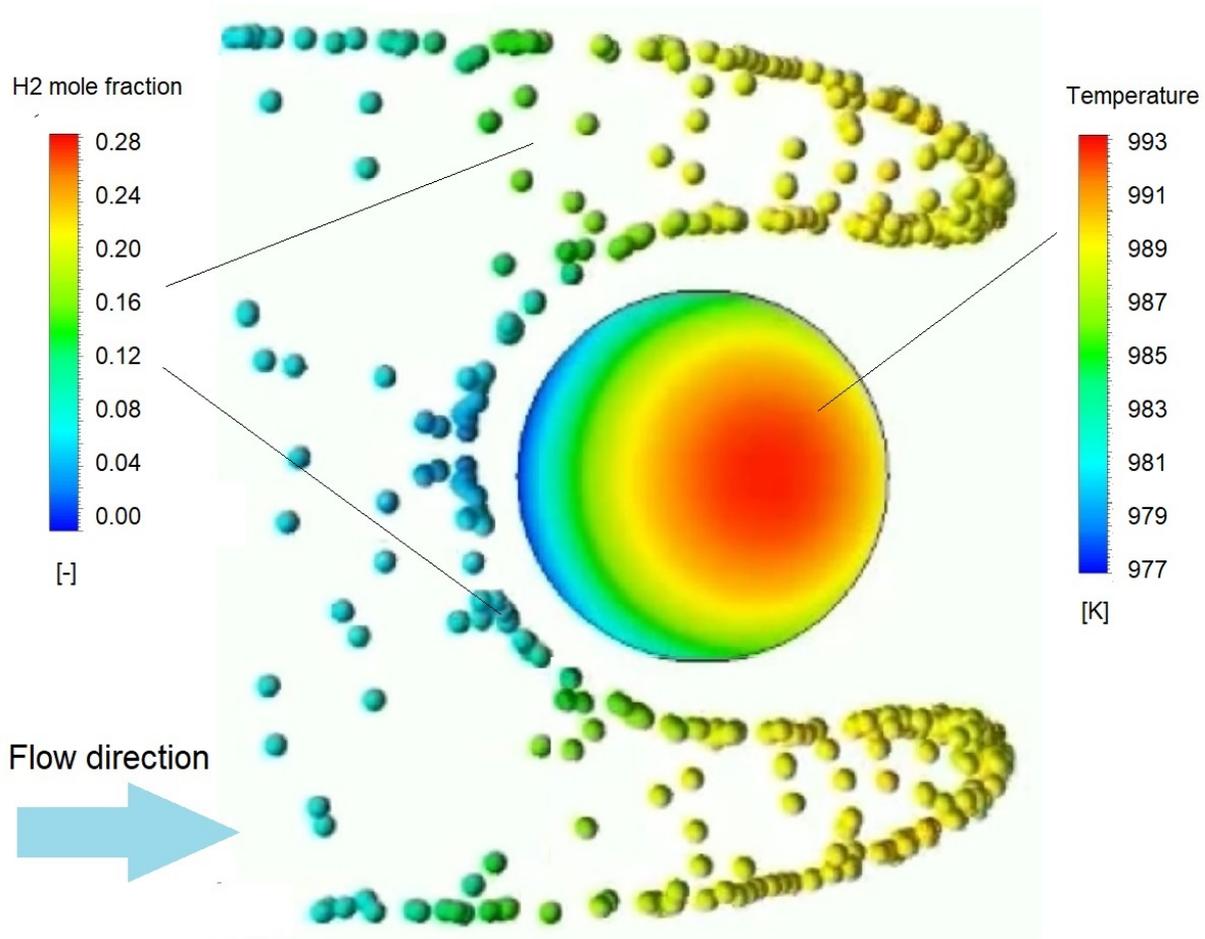
Catalyst properties	Value
Nickel content (wt.%)	9.8
S content (wt.%)	4.9
Alumina content (wt.%)	85.3
Surface area (m <sup>2</sup> /g)	155
Size of the sphere (mm)	40
Density (kg/m <sup>3</sup> )	3960
Specific heat (J/kg·K)	880
Thermal conductivity (W/m·K)	33.15
Molecular Weight (kg/kgmol)	184.41

**TABLE 2** Initial conditions

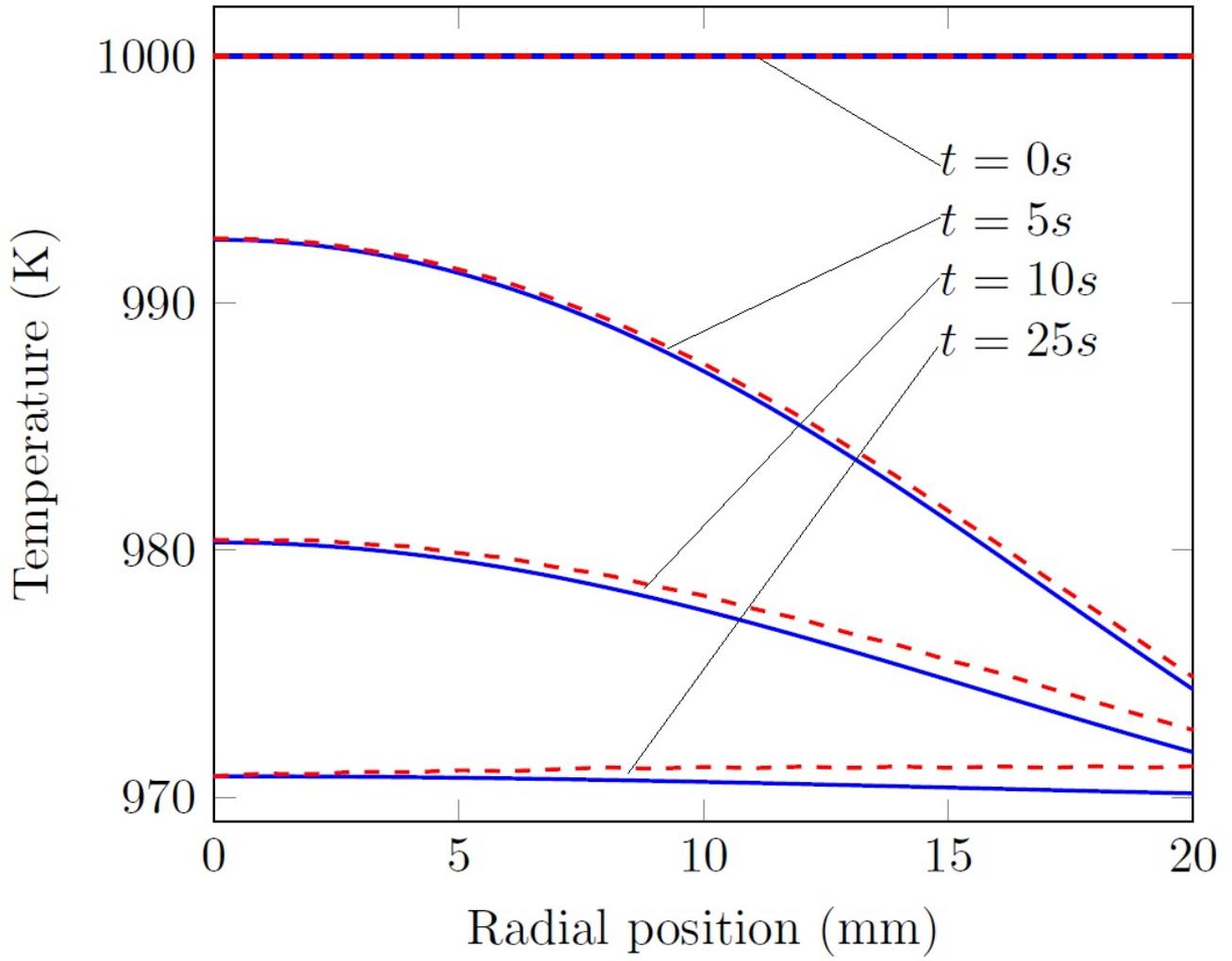
No.	Parameter	Inlet	Outlet
1	Steam-to-methane ratio	var	–
2	Temperature	var	–
3	Turbulent intensity	5 %	5 %
4	Hydraulic diameter	0.08 m	0.08 m
6	Gauge pressure	1.5 bar	0.0 bar



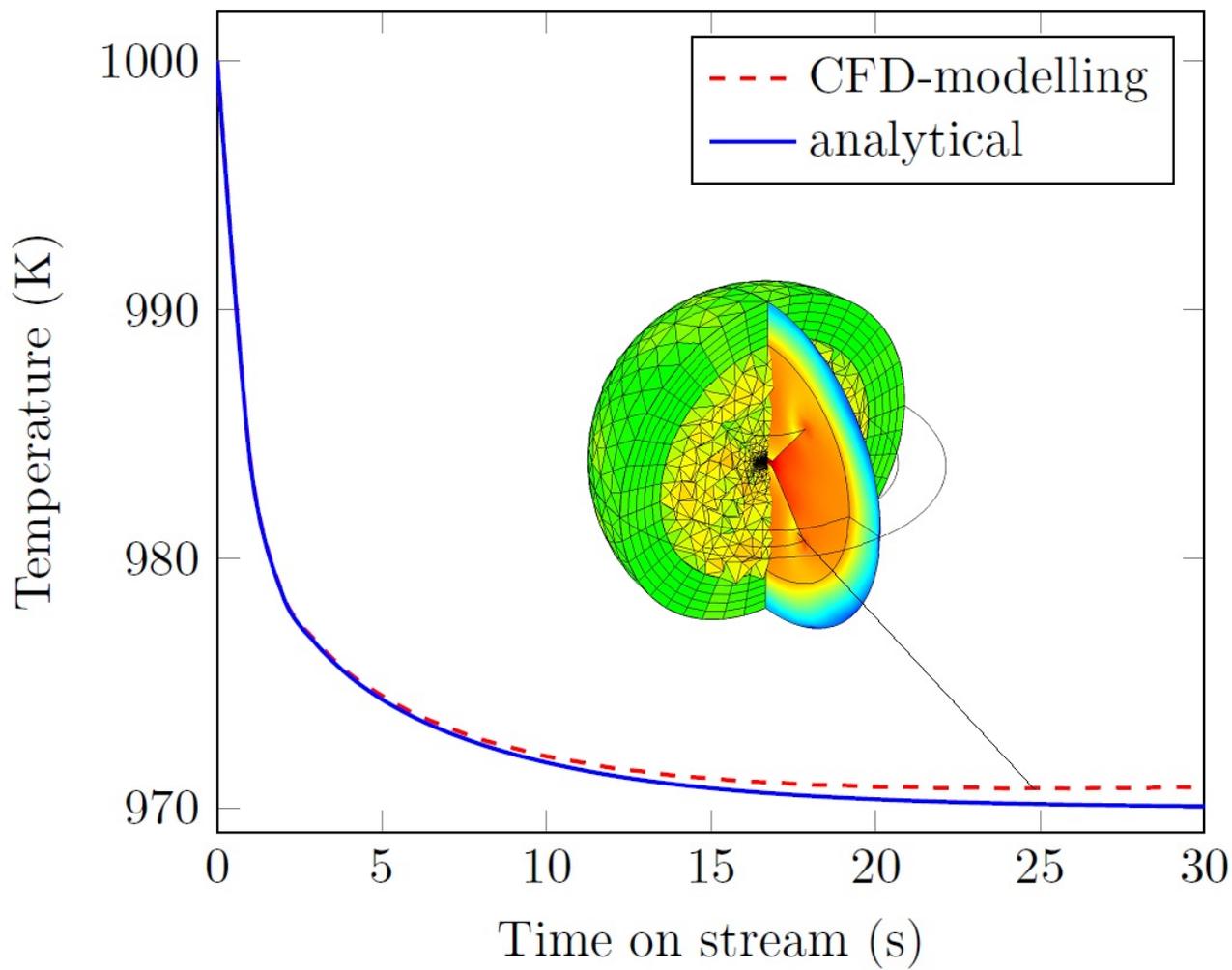
**FIGURE 1** Computational domain for CFD-modeling and schematic diagram of the steam methane reforming process.



**FIGURE 2** The temperature contour inside the catalyst particle and  $H_2$  mole fraction in the reaction space.



**FIGURE 3** The temperature profiles inside the catalyst particle: solid line – analytical results; dashed line – results of CFD-modeling.



**FIGURE 4** The temperature on the catalyst surface as a function of time on stream.

